

## Introductory Chemistry

Chemistry 110

Section: D100

Term: 2004 Fall

Instructor: Lecturer: SESSIONAL (TBA)

Discussion Topics: General Course Description:

General fundamental concepts and nomenclature; stoichiometry and chemical calculations; nuclear and atomic structures, the periodic table, the chemical bond; the properties of gases, liquids, solids and solutions; chemical kinetics and chemical equilibrium. This course has the same lecture component as Chem 111 but no laboratory work.

3 hours/week lecture; 1 hour/week tutorial; 0 hours/week laboratory.

Lectures

Topics

1 Units, measurements, Significant Figures

3 Basic concepts; Matter, Atoms, Nucleus

3 Electronic structure, the Periodic Table

1 Chemical Nomenclature

Midterm I

4 The Mole, Chemical Formula Reactions and Equations

3 Stoichiometry

3 Gas Laws

## Introductory Chemistry

2 Solutions, concentration and molarity

Midterm II

3 Ionic and Covalent Bonding

4 Acids and Bases

3 Oxidation and Reduction

3 Reaction Rates and Chemical Equilibrium

Grading: Problem Sets 10%; 2 Mid-Term Exams 40%; Final Exam 50%.

Required Texts: Stoker, "Introduction to Chemical Principles", 7th Ed. 2002, Prentice Hall.

Recommended Texts: None

Materials/Supplies: None

Prerequisite/Corequisite: Prerequisite: B.C. High School Math 12(or equiv.)(or Math 100, corequisite) or permission of the department. No previous training in chemistry is required for this course. Students with credit for high school Chem 12(or equiv.) or any university c

Notes: Students who have successfully completed B.C. High School Chem 12 (or equivalent) normally start with CHEM 121.

This outline is derived from a course outline repository database that was maintained by SFU Student Services and the University's IT Services Department. The database was retired in 2014 and the data migrated to SFU Archives in 2015.